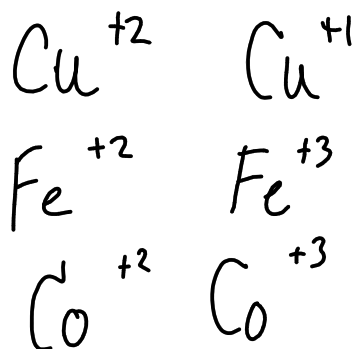
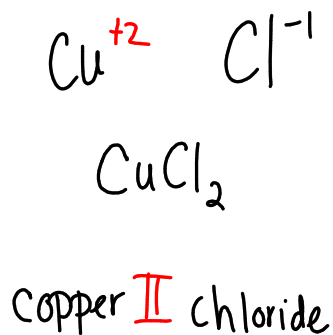
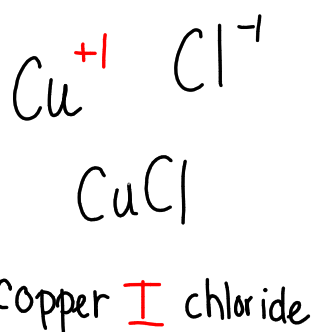


Transition Metals

Ionic compounds involving transition metals are different because transition metals can have **more than one charge!**



Feb 27-7:54 AM



They have the same first name (metal) and last name (nonmetal).

Tell them apart by adding a middle name (the charge of the metal).

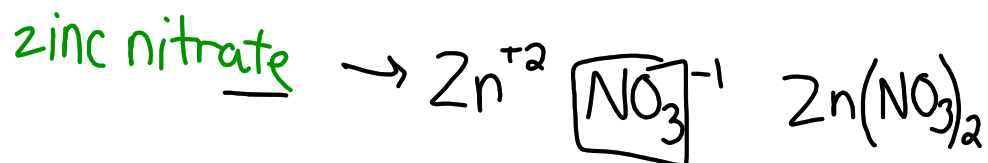
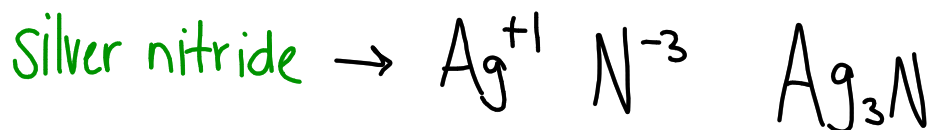
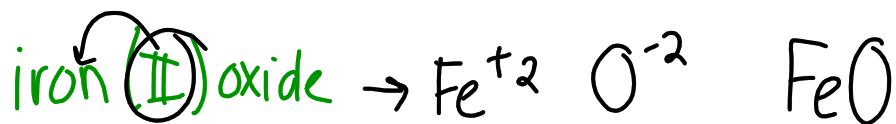
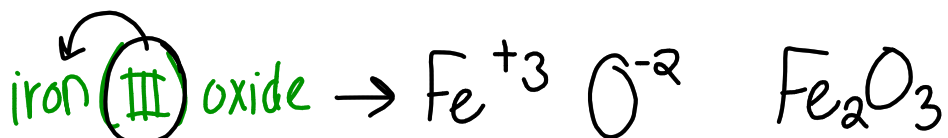
Feb 27-7:59 AM

Notes - Transition Metals

Write the charge of the metal as a Roman Numeral!

1	I	6	VI
2	II	7	VII
3	III	8	VIII
4	IV	9	IX
5	V	10	X

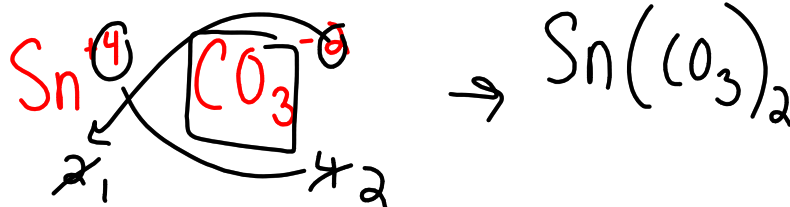
Feb 27-8:05 AM



Feb 27-8:07 AM

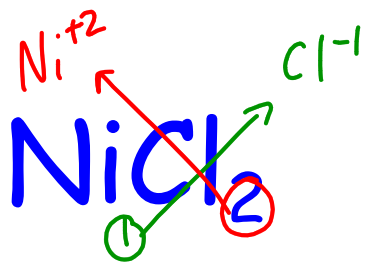
Notes - Transition Metals

tin (IV) carbonate



lead (II) sulfate

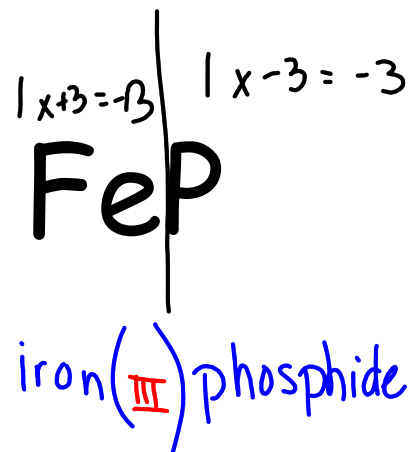
Feb 27-8:29 AM



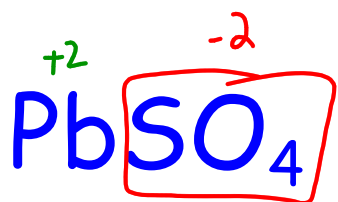
nickel (II) chloride

Mar 16-8:46 AM

Notes - Transition Metals



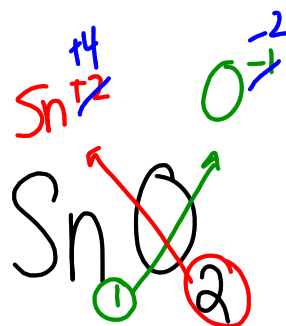
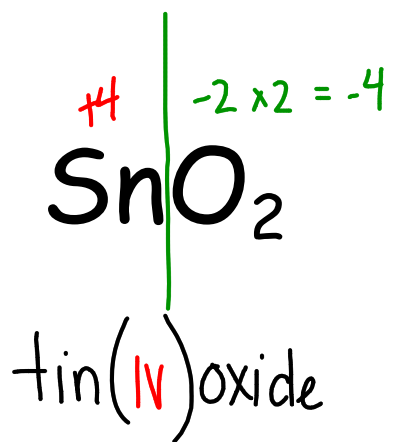
Mar 16-8:52 AM



lead II sulfate

Mar 16-8:59 AM

Notes - Transition Metals



Feb 9-8:22 AM