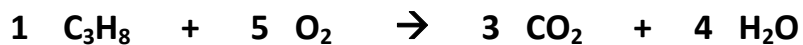


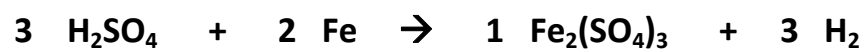
Name: _____ Date: _____ Block: _____

Stoichiometry Conversions Practice

Directions: Answer the questions below using the stoichiometry road map. You MUST SHOW ALL WORK.



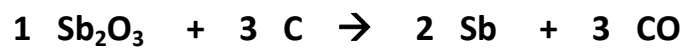
1. How many moles of CO_2 are produced from 3.2 moles of C_3H_8 ?
2. How many liters of CO_2 can be produced from 13.4 liters of oxygen gas?
3. If I have 34.6 grams of H_2O , how many grams of O_2 did I use?



4. How many moles of H_2 are made from 6.6 moles of H_2SO_4 ?
5. If given 95 grams of Fe, how many moles of $\text{Fe}_2(\text{SO}_4)_3$ are produced?
6. How many grams of H_2 are produced from 227 grams of H_2SO_4 ?



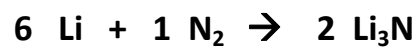
- Determine the mass in grams of LiOH produced when 0.94 grams of Li_3N reacts in the equation above.
- How many liters of water are needed to produce with 7.39 moles of NH_3 ?
- How many moles of Li_3N are needed to produce 63.8 grams of NH_3 ?



10. How many moles of carbon are needed to produce 13.9 moles of antimony?

11. If I have 50 grams of carbon then how many grams of Sb_2O_3 are needed to react?

12. Determine the moles of CO produced when 1.4 grams of Sb_2O_3 are present.



13. How many grams of nitrogen gas (N_2) are needed to completely react with 67.3 g Li?

14. If I have 3.4 moles of Li_3N then how many moles of lithium did I use?

15. How many grams of Li_3N are produced from 6.8 moles of nitrogen gas (N_2)?