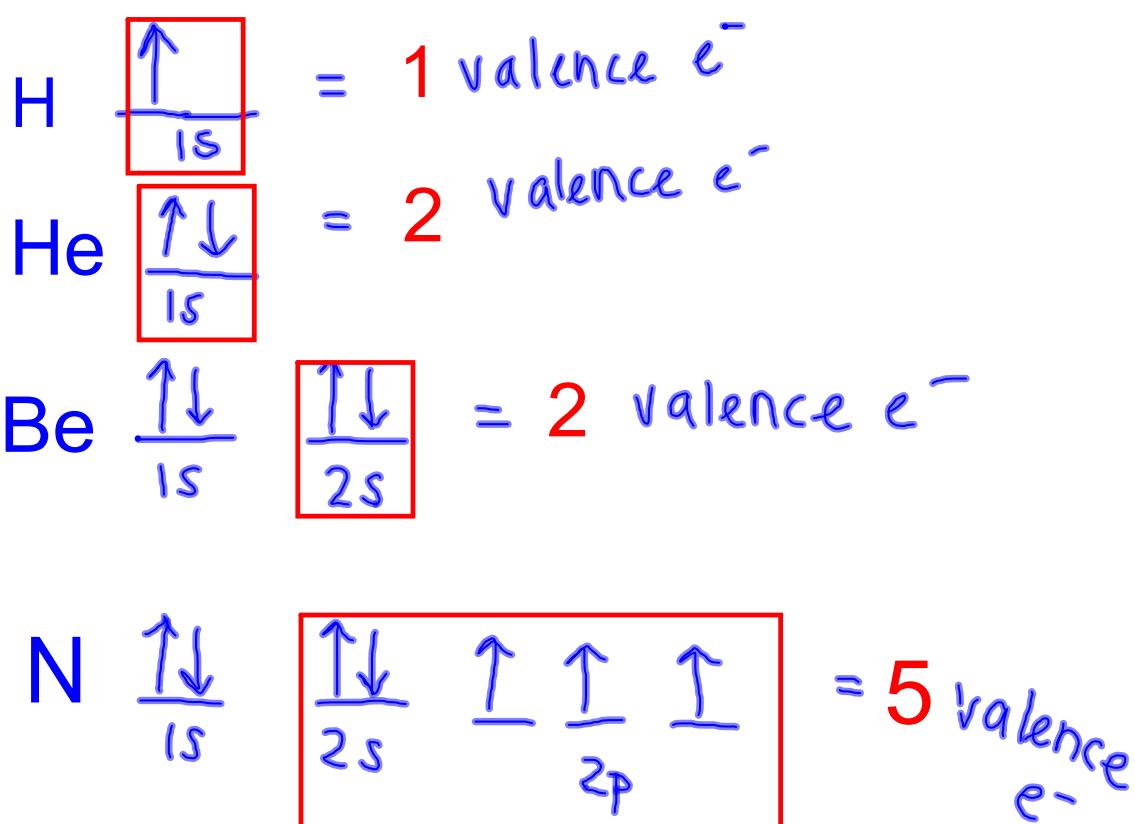


Valence Electrons:

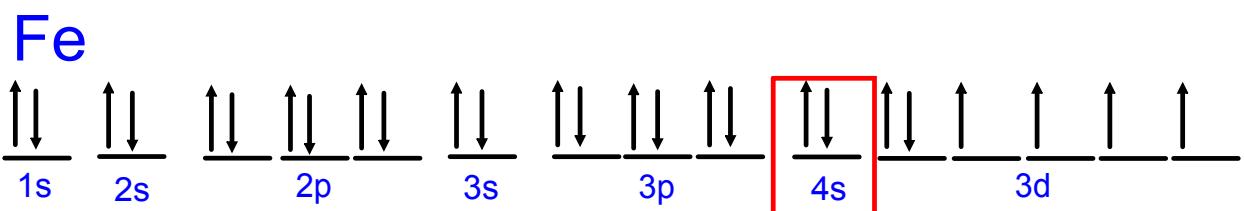
- electrons in outermost energy level in an atom
- allow an atom to bond to another to form a compound



column / group # = # of Valence e-
(look at "A" column)

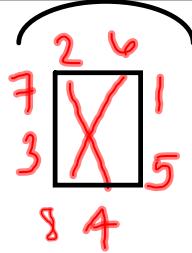
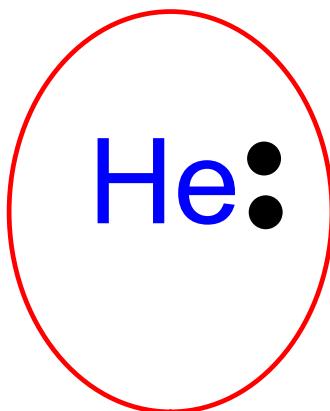
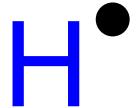
8 Valence e-maximum

2 e- from the s
6 e- from the p

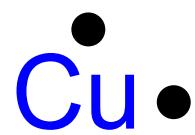
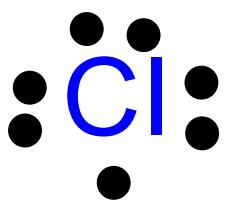
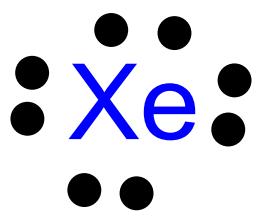
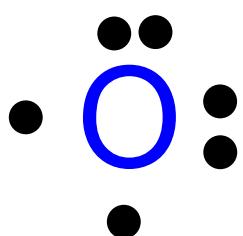
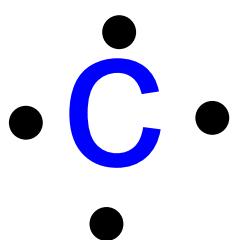


All d-block elements have
2 valence electrons because of
either the 4s, 5s, 6s, or 7s

Lewis Dot Structures:



special



Ir

Pb

Ne

Fe

P

Cs

Zn

I

Sb

Ga