

Isotope Notation

Block: _____

1. Uranium-235 and uranium-238 are considered isotopes of one another. How are uranium-235 similar, and how are they different?

2. Define isotope:

3. The number of protons in an atom is known as the _____ of that atom.
4. The number of _____ determines the name of the atom.
5. The mass number of an atom is the number of _____ plus the number of _____ in the nucleus of the atom.
6. The isotope notation for nitrogen-15 is as follows:
 - a. The number 15 is the _____ number.
 - b. The number 7 is the _____ number.
 - c. How many neutrons does nitrogen-15 have? _____
7. Write the following in isotope notation:
 - a. zinc-66: _____
 - b. chlorine-35: _____
 - c. plutonium-239: _____
 - d. helium-4: _____
 - e. uranium-235: _____
 - f. potassium-40: _____
 - g. silver-108: _____
 - h. thorium-234: _____
 - i. oxygen-16: _____
 - j. the atom with 12 protons and 12 neutrons _____
 - k. the atom with 6 protons and 7 neutrons _____
 - l. the atom with 79 protons and 117 neutrons. _____
 - m. the phosphorus atom that has 17 neutrons. _____
 - n. the copper atom that has 34 neutrons. _____
 - o. the iodine atom that has 72 neutrons _____